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- Stoichiometry Chapter 12

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Answer:  $4.93 \times 10^5$  L or 49.3 L In

Example 12.2.1 and Example

12.2.2, the identity of the limiting reactant has been apparent:

[Au(CN)<sub>2</sub>]<sup>-</sup>, LaCl<sub>3</sub>, ethanol, and

para-nitrophenol. When the limiting reactant is not apparent, we can

determine Chapter 12

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formula units KCl and three molecules O<sub>2</sub>. b Chapter 12 stoichiometry test b answer key. Four molecules NH<sub>3</sub> react with six molecules NO to form five mol- Chapter 12 Stoichiometry Answer Key Chemistry Chapter 9 Stoichiometry Test Review. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Valerie\_a\_ Key Concepts: Terms in this set (38) The efficiency of a reaction is measured by the. Percent yield. The calculated amount of product in a reaction is called the ... Theoretical yield is the answer to a ... Chemistry Chapter 9 Stoichiometry Test Review Flashcards ... CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work

Answer Key

in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub>

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Potassium oxalate 8. You will be setting up a balanced equation. Using the Law of Conservation of mass (FIRST!) you will be using Mole stoichiometry to determine the balanced equation. Unit 3

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HORROR LITERARY FICTION NON-  
FICTION SCIENCE FICTION