

Chapter 11 Review Gases Section 2

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Chapter 11 Review Gases

Section SECTION 1 Date CHAPTER

11 REVIEW Gases Class SHORT

ANSWER Answer the following

questions in the space provided. b

Pressure — orce For a constant

force, when the surface area is

tripled the surface area pressure is

(a) doubled. as much. (c) ripled. 7-0

(d) unchanged. Rank the following

pressures in increasing order. (c) 76

torr (a) 50 kPa O, OOìctbv-x Home -

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11 REVIEW Gases SECTION 2

SHORT ANSWER Answer the

following questions in the space

provided. 1. State whether the

pressure of a fixed mass of gas will

increase, decrease, or stay the

same in the following

Section 2

circumstances: increase a. temperature increases, volume stays the same decrease b. volume increases, temperature stays the same mc06se cFMSr i-vi -

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Class SHORT ANSWER Answer the

following questions in the space provided. b Pressure — orce For a constant force, when the surface area is tripled the surface area pressure is (a) doubled. as much. (c ripld. 7-0 (d) unchanged. Rank the following pressures in increasing order. Chapter 11 Review Gases Section 4 Answers Chapter 11 Gases. Chapter11. Gases. 459. t's Monday morning, and Lilia is walking out of the chemistry building, thinking about the

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introductory lecture on gases that her instructor just presented.

Dr. Chapter 11 Gases - An Introduction to Chemistry Chapter 11 Test Review. multiple choice (25) definition & applications of pressure (also atmospheric) SI unit of force. definition & use of a barometer. standard temperature & pressure (STP) Definition of Dalton's law of partial pressures. Definitions & formulas for Boyle's, Charles', Gay-Lussac's, and combined gas laws Modern Chemistry Chapter 11 GASES Chapter 11 - Gases - An Introduction to Chemistry 182 Study Guide for An Introduction to Chemistry Section Goals and Introductions ... Section 11.3 Equation Stoichiometry and Ideal Gases ... Chemistry Section 11 Answers - Free PDF File

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Quizlet CHAPTER 11 REVIEW.

Gases. SHORT ANSWER. Answer the following 9 questions in the space provided. 1. c The molar mass of a

gas at STP is the density of that gas (a) multiplied by the mass of 1 mol.

(c) multiplied by 22.4. L. CHAPTER REVIEW Gases Hydrogen molecule

1 mol H₂ at STP = 22.4 L Oxygen

molecule 1 mol O₂ at STP = 22.4 L

Carbon dioxide molecule 1 mol CO₂

at STP = 22.4 L CHAPTER 11

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