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conditions at the beginning of combustion, and  $p_4$ ,  $\rho_4$ ,  
and  $T_4$  denote conditions at the end of combustion.  
Since the volume is constant, and the mass of the gas  
is constant, then  $p_4 = \rho_3 = 11.3 \text{ kg/m}^3$ . Thus, from the  
equation of state,  $p_4 = \rho_4 RT_4 = (11.3)(287)(4000) =$   
 $1.3 \times 10^7 \text{ N/m}^2$  Anderson 7e SM - Solution manual  
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